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General Information . Molar gas

volume = $24.0 \text{ dm}^3 \text{ mol}^{-1}$ at room temperature and pressure, RTP

Avogadro constant, $N_A = 6.02 \times 10$

23 mol^{-1} Specific heat capacity of

water, $c = 4.18 \text{ J g}^{-1} \text{ K}^{-1}$ Ionic

product of water, $K_w = 1.00 \times 10^{-14}$

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mol² dm⁶ at 298 K 1 tonne = 10⁶ g
Arrhenius equation: $k = Ae^{-E_a/RT}$ or
 $\ln k = -E$

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Sheet - Beal High School*

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